Readorium Alignment to FOSS Kit: Interactions of Matter		
Readorium Books By Standard	**Magazine Articles (A) and Science Alive Videos (V) By Standard**	**Teacher Resource Center Classroom Strategy Lessons (CL) with Articles (A) by Standard**
NGSS: 6-8-PS1.A. Matter and Its Interactions: Structure and Properties of matter: Each pure substance has characteristic physical and chemical properties (for any bulk quantity under given conditions) that can be used to identify it. (MS-PS1-2), (MS-PS1-3) Gases and liquids are made of molecules or inert atoms that are moving about relative to each other. (MS-PS1-4) In a liquid, the molecules are constantly in contact with others; in a gas, they are widely spaced except when they happen to collide. In a solid, atoms are closely spaced and may vibrate in position but do not change relative locations. (MS-PS1-4)		
• Chemical and Physical Properties of Matter 1 • Chemical and Physical Properties of Matter 2 • Formation of Volcanoes • Lights Sound Action • Plate Tectonics • Weather • Pollution	• Crime Scene Science(A) • Matter Matters(A) • Splash(A) • The Water Cycle(A)	• Determining Importance (CL-3, A-2 Crystals)
NGSS: 6-8-PS1.B. Matter and Its Interactions: Chemical Reactions: Substances react chemically in characteristic ways. In a chemical process, the atoms that make up the original substances are regrouped into different molecules, and these new substances have different properties from those of the reactants. (MS-PS1-2), (MS-PS1-3), (MS-PS1-5)		
• Chemical and Physical Properties of Matter 1 • Chemical and Physical Properties of Matter 2 • The Formation of Volcanoes	• Cafeteria Chemistry: How to Play with Your Food and Astound Your Friends (A) • Crystals (A) • Gold - The Magnificent Metal (A) • Kitchen Chemistry (A) • The Cool World of Chemistry (A) • Excuse me, but burping is natural(A) • The Science of Movie Stunts(A)	• Creating Sensory Images (CL-1, A-2 Kitchen Chemistry) • Determining Importance (CL-3, A-2 Crystals)
NGSS: 6-8-PS3.A. Energy: Definitions of Energy: Temperature is a measure of the average kinetic energy of particles of matter. The relationship between the temperature and the total energy of a system depends on the types, states, and amounts of matter present. (MS-PS3-3), (MS-PS3-4) The term “heat” as used in everyday language refers both to thermal motion (the motion of atoms or molecules within a substance) and radiation (particularly infrared and light). In science, heat is used only for this second meaning; it refers to energy transferred when two objects or systems are at different temperatures. (secondary to MS-PS1-4) Temperature is not a measure of energy; the relationship between the temperature and the total energy of a system depends on the types, states and amounts of matter present. (secondary to MS-PS1-4)		
• Lights Sound Action • Sports Physics • Newton’s Laws	• Weapons Older than Dirt: The History of Some of the World's Most Ancient Weapons (A) • Things That Go BOOM!: The History and Chemistry of Explosives (A)	